

B. Pharm. (CBCS Pattern) Semester-I
BP102T - Pharmaceutical Analysis-I

P. Pages : 1

Time : Three Hours



GUG/W/24/10871

Max. Marks : 75

- Notes :
1. Diagrams and Chemical equation should be given wherever necessary.
 2. Illustrate your answers wherever necessary with the help of neat sketches.
 3. All questions are compulsory.

- 1.** Solve all of the following questions are compulsory. **2x10**
=20
- a) Define the term Analysis. Give its scope and application.
 - b) Define normality, molarity and equivalent weight.
 - c) What are primary standard and secondary standards.
 - d) Explain principle of limit test of chloride.
 - e) What is neutralization curve? Enlist the type of neutralization curve.
 - f) What is complexometric titration? Enlist the types of metal ion indicators.
 - g) Draw labelled diagram of dropping mercury electrode.
 - h) Define Ostwald ripening. Enlist the steps involved in gravimetric analysis.
 - i) What is reference electrode? Enlist them.
 - j) Define accuracy and precision.
- 2.** Solve the following (**any two**) **10x2**
=20
- a) What are errors? Explain sources of errors and methods to minimize errors.
 - b) Explain in detail steps involved in gravimetric analysis.
 - c) Explain in detail alkalimetry and acidimetry, with types of solvent used in non aqueous titration.
- 3.** Solve the following (**any seven**) **5x7**
=35
- a) What are the different methods to express concentration.
 - b) Explain about limit test for arsenic.
 - c) Explain different theories of acid and base.
 - d) Explain diazotization titration in short.
 - e) Give an account on types of complexometric titration.
 - f) Explain Mohr's method with estimation of sodium chloride.
 - g) Explain iodimetry and iodometry.
 - h) Write note on types of conductometric titration.
 - i) Explain types of indicator electrode used in potentiometry.
