

B.Sc.-II (CBCS Pattern) Semester - III
USCCHT06 - Chemistry Paper-II : Physical Chemistry

P. Pages : 2

Time : Three Hours



GUG/S/23/11601

Max. Marks : 50

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- Notes : 1. All **five** questions are compulsory and carry equal marks.
2. Draw Diagram wherever necessary.
3. Use of calculator is permitted.

1. a) Draw the phase diagram for lead-silver system. Discuss the application of phase rule to this system. 5
- b) Discuss phenol-water system of partially miscible liquids. Explain the effect of impurities on critical solution temperature. 5

OR

- c) State phase Rule. Explain the meaning of terms involved in it. 2½
- d) Discuss partial miscibility of nicotine – water system. 2½
- e) Discuss minimum boiling azeotropes ethanol water system. 2½
- f) Explain deviation of Nerst Distribution law when solute undergoes dissociation in one of the solvents. 2½
2. a) Derive an equation for entropy change for an ideal gas in term of Pressure (P) & temperature (T). 5
Calculate the entropy change when 2 moles of an ideal gas is allowed to expand at 300K from pressure of 10 atm to 2 atm.
- b) Derive an integrated form of Van't Hoff reaction isochore. 5

OR

- c) What are the needs for second law of thermodynamics? 2½
- d) Derive Gibb's – Helmholtz equation. 2½
- e) Derive the relation between standard free energy change and equilibrium constant. 2½
- f) State & explain chemical potential. 2½
3. a) What is first order reaction? Derive an expression for rate constant of first order reactions. 5
- b) Derive expression for rate constant based on equilibrium constant and thermodynamic aspects. 5

OR

- c) Discuss effect of temperature on rate of reaction. 2½
- d) The rate constant of reaction is $2.5 \times 10^{-4} \text{ Sec}^{-1}$ at 35°C and $3.9 \times 10^{-4} \text{ sec}^{-1}$ at 45°C . Calculate energy of activation of reaction ($R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$). 2½
- e) Discuss the characteristics of catalyzed reaction. 2½
- f) Write a note on enzyme catalysis. 2½
4. a) What are elevation of boiling point? Obtain an expression for molar mass of substance from elevation of boiling point. 5
- b) Discuss determination of magnetic susceptibility by using Gouy method. 5

OR

- c) Explain diamagnetism & paramagnetism with suitable examples. 2½
- d) State Raoult's Law of lowering of vapour pressure. How can it be used to determine the molecular weight of non-volatile solute in solution. 2½
- e) 0.440 gm of substance dissolved in 22.2 g of benzene lowered the freezing point of benzene by 0.567°C . Calculate the molar mass of substance. ($K_f = 5.12^\circ\text{C mol}^{-1}$). 2½
- f) Define degree of association and obtain the relation between degree of association and Van't Hoff factor. 2½
5. Attempt **any ten**. **1x10**
=10
- i) State Henry's law and give any one limitation.
- ii) Write Clausius – Clapeyron equation in its integrated form.
- iii) Define lower and upper Consolute temperature.
- iv) Write the statement of second law of thermodynamics.
- v) What do you mean by partial molar quantities?
- vi) Define standard free energy.
- vii) What is zero order reaction?
- viii) Define.
- a) Molecularity of reaction b) Half life of a reaction
- ix) What is heterogeneous catalytic reaction?
- x) What do you mean by osmotic pressure?
- xi) What is Van't Hoff factor?
- xii) Calculate magnetic moment of Mn^{+2} .
