

Max. Marks : 75

P.T.O

- 10) Non – aqueous titration is carried out for -----
 - a) Water insoluble drugs
 - b) Weakly acidic drugs
 - c) Weakly basic drugs
 - d) All of the above
- 11) Solvent effect is also called as
 - a) Levelling effect
 - b) Leasing effect
 - c) Additive effect
 - d) None of the above
- 12) End point in Mohr's method is indicated by -----
 - a) Red Colour
 - b) Blue Colour
 - c) Yellow Colour
 - d) Black Colour
- 13) In Volhard's method, the solution filled in to the burette is -----
 - a) Silver Nitrate
 - b) Ammonium Thiocyanate
 - c) Ferric Thiocyanate
 - d) Potassium Chromate
- 14) Types of masking techniques used in Complexometric Titrations are -----
 - a) Masking by precipitation
 - b) Masking by oxidation and reduction
 - c) Mixed and kinetic masking
 - d) All of the above
- 15) In Alkalimetric titration the titrant is -----
 - a) Iodin
 - b) EDTA
 - c) NaOH
 - d) Magnesium
- 16) Following are properties of good precipitate -----
 - a) Easily filtered and washed free of contaminants
 - b) Unreactive with constituents of the atmosphere
 - c) (a) and (b)
 - d) None of the above
- 17) Analyte which produces color change in redox reaction at end point can be used as -----
 - a) Self indicator
 - b) Starch indicator
 - c) Alkalimetric indicator
 - d) Both (a) and (c)
- 18) The electrode which indicates redox potential in redox titration is -----
 - a) Combination electrode
 - b) Reference electrode
 - c) Indicator electrode
 - d) All of the above
- 19) Factors affecting conductance are -----
 - a) Dilution
 - b) Temperature
 - c) Types of ion
 - d) All of the above
- 20) Which of the following is not a reference electrode?
 - a) Silver – silver chloride electrode
 - b) ISFET
 - c) SHE or NHE
 - d) Calomel electrode

2. Attempt **any two** out of three long answer questions.

**2x10
=20**

- a) What is Gravimetric analysis? Explain the steps involved in Gravimetric Analysis.

- b) What is complexometric titration? Write in short the factors affecting stability of complexes. Explain the types of EDTA titration.
- c) Write the end point detection in Aqueous and non – aqueous acid base titration with examples. And write about advantages and array by non – aqueous titration.

3. Attempt **any seven** out of nine questions.

**7x5
=35**

- a) Write about Iodometry and Iodimetry.
- b) Write note on Volhard's method.
- c) Explain about the reference electrode in Conductometry titration.
- d) Explain in brief about the significant figure.
- e) Write note on Amphoteric and protogenic solvent.
- f) What is Arrhenius and Bronsted Lowry concept.
- g) Write note on applications of diazotization titrations.
- h) Explain theory of redox titrations.
- i) Explain conductometric titrations and it's principle.
