B.Pharm. First Year CBCS Pattern Semester-I

BP 102T - Pharmaceutical Analysis-I

P. Pages : 3 **GUG/W/23/10871**

Time : Three Hours * 5 9 6 8 * Max. Marks : 75

Notes : 1. All questions are compulsory.
2. All questions carry as indicated marks.
3. Illustrate your answers wherever necessary with the help of neat sketches.

1. Multiple Choice Questions. **20**

i) ---------- analysis is the determination of the amount of particular element, species or compound present in sample.
a) Quantitative analysis b) Quanlitative analysis
c) Semi-quantitative d) Semi qualitative

ii) % W/W express
a) No. of gm of solute in 1000gm of product
b) No. of gm of solute in 100gm of product
c) No. of gm of solute in 1000ml of product
d) No. of gm of solute in 100ml of product

iii) Lead acetate cotton plug is used to trap
a) Hydrochloric acid b) Hydrogen Sulphide
c) Chloroform d) None of these

iv) The process of adding known concentration until it complete the reaction with known volume s called as
a) Precipitation b) Complexation
c) Titration d) None of these

v) Errors is the difference between--------&-----
a) Experimental mean and measured mean
b) Measure value and true value
c) Both a and b
d) None of these

vi) Amphiprotic solvents are----&-----
a) Aprotic & Diprotic b) Protophillic & Protogenic
c) Aprotic & Protogenic d) None of these

vii) Non aqueous titration is carried out for
a) Water insoluble drugs b) Weakly acidic drug
c) Weakly basic drug d) All of these

viii) The end point of a titration is defined as
a) The equivalence point of the titration
b) The actual measured volume of titrant required to complete a titration
c) The volume associated with actual stoichiometric quantity of titrant required to complete a titration.
d) None of these

ix) The acid base indicators exist in two tautomeric forms in
a) Ostwald theory b) Quinonoid theory
c) Both a and b d) None of these

x) Carbon tetrachloride is an example of
a) Amphiprotic solvent b) Aprotic solvent
c) Both a and b d) All of these

xi) In which method ferric ammonium sulphate is used as an indicator
a) Fajan's method b) Mohr's method
c) Volhard's method d) None of these

xii) Example of sequestering agent is
a) Nitroglycerine b) Silver nitrate
c) Lead nitrate d) EDTA

xiii) Which is used as masking agent for lead
a) Sodium Sulphide b) Oxalates
c) Thiocetamide d) All of these

xiv) Given the following are properties of good precipitate except
a) Easily filtered and washed free of contamination
b) Significant loss of the analyte occurs during filtration and washing.
c) Unreactive with constituents of the atmosphere.
d) Both a and b

xv) In diazotization titration reaction temperature maintained is
a) 0-5 b) 10-15
c) 15-25 d) None of these

xvi) In a redox titration
a) Oxidation occurs b) Reduction occurs
c) Neutralization occurs d) Both a and b

xvii) The reciprocal of resistivity is known as
a) Molar conductance b) Specific conductance
c) Equivalent conductance d) Permittivity

xviii) The value of equivalent conductance
a) Decrease with pressure b) Increase with pressure
c) Decrease with dilution d) Increase with dilution

xix) An electrode, whose electrode potential is well known and stable is a
a) Indicator electrode b) Reference electrode
c) Both a and b d) None of these

xx) For qualitative analysis by polarography, characteristic parameter used is
a) Diffusion current b) Half wave potential
c) Voltage d) None of these

2. Solve **any two**. **20**

a) What are errors? Explain methods to minimize errors in detail.

b) Explain the step involved in Gravimetric analysis in detail.

c) Explain in detail electrodes used in potentiometry.

3. Solve **any seven**. **35**

a) Explain the different methods to express concentration.

b) Write method to prepare and standardization of sodium hydroxide.

c) Explain in detail limit test for Iron.

d) Explain Mohr’s and Volhard's method.

e) Write note on masking and demasking agent. Explain estimation of magnesium sulphate

f) Write note on iodimetry and iodometry.

g) Write note on conductivity cell with applications of conductometry.

h) What are alkalimetry and acidimetry.

i) Explain construction and working of dropping mercury electrode.

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